

➤ ASSESS YOURSELF

Very Short Answer Type Questions [1 Mark]

1. Why gases are compressible but not liquids?
2. How is matter classified on the basis of physical state?
3. What do you think, matter is made up of small particles or not?
4. Why does the level of water not change when salt is dissolved in water?
5. What about a rubber band, can it change its shape? Is it solid?
6. Sugar and salt kept in different jars gain the shape of their respective jars. Are they solids?
7. What is CGS unit of volume?
8. How many cm^3 are there in 1 L?
9. What is meant by density?
10. Why is density of solids mostly higher than that of liquids?
11. A sponge can be compressed. Is it solid or not? Why?
12. What is meant by kinetic energy?
13. Why do liquids flow?

Short Answer Type Questions [I] [2 Marks]

14. Why do people in villages use earthen pots in summer to cool water?
15. Convert 100°C to kelvin scale.
16. Write the name of two substances that sublime.
17. Give two factors which increase the rate of evaporation.
18. Why does water have higher boiling point than alcohol? At what temperature in Kelvin, it can be changed into solid state and into vapours?

Short Answer Type Questions [II] [3 Marks]

19. (a) Name the term used for:
 - (i) Solid directly formed from the gas;
 - (ii) Gas directly formed from the solid.(b) Ice at 0°C is more effective in cooling than water at 0°C . Give reason.
20. Show experimentally that matter is made up of small particles.
21. Describe by experiment the relationship between rate of diffusion and density of liquid.
22. How is melting point related to intermolecular forces of attraction? Why is latent heat of vapourisation higher than latent heat of fusion?
23. Show by an experiment that by increasing the temperature, substance can be changed into liquid and then into vapours.
24. Draw a cyclic figure to show interconversion of states and explain fusion, vapourisation, condensation, solidification and sublimation.

AYAZ SIR'S CBSE COACHING
RISING SCIENCE
Matter In Our Surrounding

- Q1. What is dry ice?
- Q2. What is normal atmospheric pressure?
- Q3. Give two examples of diffusion.
- Q4. At what temperature is the density of water maximum?
- Q5. Name the factors that affect evaporation.
- Q6. With the help of an example, explain how diffusion of gases in water is essential?
- Q7. How is the high compressibility property of gas useful to us?
- Q8. Liquids and gases can be compressed but it is difficult to compress solids. Why?
- Q9. State the differences between solid, liquid and gas.
- Q10. Pressure and temperature determine the state of a substance. Explain.
- Q11. Give reasons for the following:
- (a) Water kept in earthen pots become cool during summer.
 - (b) Naphthalene balls disappear with time without leaving any solid.
 - (c) A gas fills completely the vessel in which it is kept.

Ayaz sir's CBSE Coaching

Class -X

Marks15

Is Matter Around Us Pure

1. What is meant by pure substance? 1m
2. Distinguish between homogeneous and heterogeneous substance. 2m
3. How are sol, solution different from each other? 2m
4. How will you separate a mixture containing kerosene and petrol which are miscible with each other? 2m
5. Write a note on sublimation. 2m
6. What is chromatography? What are its advantages over other methods of separation? 2m
7. Describe the process of filtration. 2m
8. Why does the beam of light illuminated when it enters a room through a small hole? 2m

www.ayazsir.com